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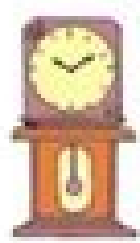
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Name: _____

Elapsed Time: to the quarter hour

How much time has passed?
Write the elapsed time.



Start	End	Elapsed Time
3:00 P.M.	to 11:45 P.M.	____hour(s) ____minutes
1:15 P.M.	to 2:00 P.M.	____hour(s) ____minutes
2:30 P.M.	to 8:45 P.M.	____hour(s) ____minutes
6:45 P.M.	to 10:30 P.M.	____hour(s) ____minutes
9:30 P.M.	to 12:15 A.M.	____hour(s) ____minutes
8:45 A.M.	to 12:45 P.M.	____hour(s) ____minutes
4:00 A.M.	to 11:15 A.M.	____hour(s) ____minutes
10:30 A.M.	to 2:45 P.M.	____hour(s) ____minutes
7:00 A.M.	to 10:30 A.M.	____hour(s) ____minutes
12:00 A.M.	to 10:45 A.M.	____hour(s) ____minutes
*12:00 P.M	to 12:45 A.M.	____hour(s) ____minutes

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The image shows several pages of math worksheets. The top page features multiplication and division problems. The middle page has a word problem about a school bus. The bottom page is a grid-in response area for a math test.

For example, water, H₂O, and hydrogen peroxide, H₂O₂, are alike in that their respective molecules are composed of hydrogen and oxygen atoms. Explore more about the mole by reviewing the information under "Dig Deeper." The relationships between formula mass, the mole, and Avogadro's number can be applied to compute various quantities that describe the composition of substances and compounds. Referring to the periodic table, the atomic mass of K is 39.10 amu, and so its molar mass is 39.10 g/mol. Analysis of these neurons from mice subjected to various drug therapies shows significant changes in the average number of dopamine molecules contained in individual vesicles, increasing or decreasing by up to three-fold, depending on the specific drug used. Consider chloroform (CHCl₃), a covalent compound once used as a surgical anesthetic and now primarily used in the production of the "anti-stick" polymer, Teflon. Performing the calculation, we get: $[1.42 \times 10^{-4} \times (1.75 \times 10^{-1} \times 12.01 + 3 \times 35.45)] \times 6.022 \times 10^{23}$. This is consistent with the anticipated result. Following the approach described above, the average molecular mass for this compound is therefore: $(0.1 \times 12.01 + 3 \times 35.45) \text{ g/mol} = 119.37 \text{ g/mol}$. Today, we possess sophisticated instruments that allow the direct measurement of these defining microscopic traits; however, the same traits were originally derived from the measurement of macroscopic properties (the masses and volumes of bulk quantities of matter) using relatively simple tools (balances and volumetric glassware). Even though a sodium cation has a slightly smaller mass than a sodium atom (since it is missing an electron), this difference will be offset by the fact that a chloride anion is slightly more massive than a chloride atom (due to the extra electron). From left to right (bottom row): 32.1 g sulfur, 207 g lead, and 118.7 g tin. Ionic compounds are composed of discrete cations and anions combined in ratios to yield electrically neutral bulk matter. The few exceptions to this guideline are very light ions derived from elements with precisely known atomic masses. Likewise, the molecular mass of an aspirin molecule, C₉H₈O₄, is the sum of the atomic masses of nine carbon atoms, eight hydrogen atoms, and four oxygen atoms, which amounts to 180.15 amu (Figure 2). However, because a hydrogen peroxide molecule contains two oxygen atoms, as opposed to the water molecule, which has only one, the two substances exhibit very different properties. The molecule shows the molecular structure of chloroform. The mole provides a link between an easily measured macroscopic property, number of atoms, molecules, and so forth. Keep in mind, however, that the formula for an ionic compound does not represent the composition of a discrete molecule, so it may not correctly be referred to as the "molecular mass." As an example, consider sodium chloride, NaCl, the chemical name for common table salt. Since the amount of Ar is less than 1 mole, the mass will be less than 1 mole of Ar, approximately 40 g. Solution As for elements, the mass of a compound can be derived from its molar amount as shown: The molar mass for this compound is computed to be 176.124 g/mol. The identity of a substance is defined not only by the types of atoms or ions it contains, but by the quantity of each type of atom or ion. Figure 1. A mole is defined as the amount of substance containing the same number of discrete entities (such as atoms, molecules, and ions) as the number of atoms in a sample of pure 12C weighing exactly 12 g. The average mass of a chloroform molecule, CHCl₃, is 119.37 amu, which is the sum of the average atomic masses of each of its constituent atoms. The molecular formula of chloroform indicates that a single molecule contains one carbon atom, one hydrogen atom, and three chlorine atoms. One Latin connotation for the word "mole" is "large mass" or "bulk," which is consistent with its use as the name for this unit. Dividing the compound's mass by its molar mass yields: $(183.18 \text{ g}) / (119.37 \text{ g/mol}) = 1.53 \times 10^{-3} \text{ mol}$. The compound's formula shows that each molecule contains seven carbon atoms, and so the number of C atoms in the provided sample is: $(1.31 \times 10^{-2} \text{ mol}) \times (7 \text{ C atoms/molecule}) \times (6.022 \times 10^{23} \text{ molecules/mol}) = 5.6 \times 10^{22} \text{ C atoms}$. In this case, logic dictates (and the factor-label method supports) multiplying the provided amount (mol) by the molar mass (g/mol): $(1.31 \times 10^{-2} \text{ mol}) \times (176.124 \text{ g/mol}) = 2.31 \text{ g}$. The given number of moles is a very small fraction of a mole (~10⁻³), and so the corresponding mass should be roughly one-thousandth of the molar mass (~0.04 g). In this case, logic dictates (and the factor-label method supports) multiplying the provided amount (mol) by the molar mass (g/mol): $(1.31 \times 10^{-2} \text{ mol}) \times (176.124 \text{ g/mol}) = 2.31 \text{ g}$. The result is in agreement with our expectations, around 0.04 g Ar. Check Your Learning What is the mass of 2.561 mol of gold? For purposes of computing a formula mass, it is helpful to rewrite the formula in the simpler format, Al₂(SO₄)₃. One mole of glycine, C₂H₅O₂N, contains 2 moles of carbon, 5 moles of hydrogen, 2 moles of oxygen, and 1 mole of nitrogen: The provided mass of glycine (~28 g) is a bit more than one-third the molar mass (~75 g/mol), so we would expect the computed result to be a bit greater than one-third of a mole (~0.33 mol). Figure 8. (a) A typical human brain weighs about 1.5 kg and occupies a volume of roughly 1.1 L. Deriving Grams from Moles for a Compound A liter of air contains 9.2 × 10⁻⁴ mol argon. Solution Moles of this compound are comprised of 13 carbon atoms, 18 hydrogen atoms, and 2 oxygen atoms. This relationship holds for all elements, since their atomic masses are measured relative to that of the amu-reference substance, 12C. Consistent with its definition as an amount unit, 1 mole of any element contains the same number of atoms as 1 mole of any other element. Moreover, the mass of an electron is negligibly small with respect to the mass of a typical atom. The number of molecules in a single droplet of water is roughly 100 billion times greater than the number of people on earth. According to the definition of the mole, 12 g of 12C contains 1 mole of 12C atoms (its molar mass is 12 g/mol). Sodium chloride is an ionic compound composed of sodium cations, Na⁺, and chloride anions, Cl⁻, combined in a 1:1 ratio. (credit: "tanakawho"/Wikimedia commons) The mole is used in chemistry to represent 6.022 × 10²³ of something, but it can be difficult to conceptualize such a large number. The number of entities composing a mole has been experimentally determined to be 6.02214179 × 10²³, a fundamental constant named Avogadro's number (NA) or the Avogadro constant in honor of Italian scientist Amedeo Avogadro. The molar amount of a substance may be calculated by dividing its mass (g) by its molar mass (g/mol). The factor-label method supports this mathematical approach since the unit "g" cancels and the answer has units of "mol": $(1.47 \text{ g}) / (0.5 \text{ g/mol}) = 2.94 \text{ mol}$. The calculated magnitude (0.12 mol K) is consistent with our ballpark expectation, since it is a bit greater than 0.1 mol. For example, if we know the mass and chemical composition of a substance, we can determine the number of moles and calculate number of atoms or molecules in the sample. Figure 7. Dopamine is involved in various neurological processes that impact a wide variety of human behaviors. What is the molecular mass (amu) for this compound? For covalent substances, the formula represents the number and types of atoms composing a single molecule of the substance; therefore, the formula mass may be correctly referred to as a molecular mass. Solution The molar amount of Ar is provided and used to derive the corresponding mass in grams. Using miniaturized probes that can selectively detect dopamine molecules in very small amounts, scientists have determined that the vesicles of a certain type of mouse brain neuron contain an average of 30,000 dopamine molecules per vesicle (about 5 × 10⁻²⁰ mol or 50 zmol). Solution The formula for this compound indicates it contains Al³⁺ and SO₄²⁻ ions combined in a 2:3 ratio. While atomic mass and molar mass are numerically equivalent, keep in mind that they are vastly different in terms of scale, as represented by the vast difference in the magnitudes of their respective units (amu versus g). Check Your Learning A prospector panning for gold in a river collects 15.00 g of pure gold. It sends and receives signals to and from muscles and other internal organs to monitor and control their functions; it processes stimuli detected by sensory organs to guide interactions with the external world; and it houses the complex physiological processes that give rise to our intellect and emotions. The formula mass for this compound is computed as 58.44 amu (see Figure 3). What is the mass of Ar in a liter of air? Solution The mass of K is provided, and the corresponding amount of K in moles is requested. The mole is an amount unit similar to familiar units like pair, dozen, gross, etc. How many copper atoms are in 5.00 g of copper wire? Deriving Moles from Grams for an Element According to nutritional guidelines from the US Department of Agriculture, the estimated average requirement for dietary potassium is 4.7 g. Computing Formula Mass for an Ionic Compound Aluminum sulfate, Al₂(SO₄)₃, is an ionic compound that is used in the manufacture of paper and in various water purification processes. One of these amino acids is glycine, which has the molecular formula C₂H₅O₂N. Check Your Learning Beryllium is a light metal used to fabricate transparent X-ray windows for medical imaging instruments. How many Au atoms are in this quantity of gold? Chemical signaling occurs at the interface between different neurons when one of the cells releases molecules (called neurotransmitters) that diffuse across the small gap between the cells (called the synapse) and bind to the surface of the other cell. What is the mass of this allowance in grams? Deriving Grams from Moles for a Compound Vitamin C is a covalent compound with the molecular formula C₆H₈O₆. Significant progress has been made recently in directly measuring the number of dopamine molecules stored in individual vesicles and the amount actually released when the vesicle undergoes exocytosis. Figure 5. Figure 6. Solution We can derive the number of moles of a compound from its mass following the same procedure we used for an element in Example 3. The molar mass of glycine is required for this calculation, and it is computed in the same fashion as its molecular mass. Extending this principle, the molar mass of a compound in grams is likewise numerically equivalent to its formula mass in amu (Figure 5). Copper wire is composed of many, many atoms of Cu. (credit: Emilian Robert Vicoli) Solution The number of Cu atoms in the wire may be conveniently derived from its mass by a two-step computation: first, calculating the molar amount of Cu, and then using Avogadro's number (NA) to convert this molar amount to number of Cu atoms. Considering that the provided sample mass (5.00 g) is a little less than one-tenth the mass of 1 mole of Cu (~64 g), a reasonable estimate for the number of atoms in the sample would be on the order of one-tenth NA, or approximately 10²² Cu atoms. Great progress has been made in brain research over the past few decades, and the BRAIN Initiative, a federal initiative announced in 2013, aims to accelerate and capitalize on these advances through the concerted efforts of various industrial, academic, and government agencies (more details available at www.whitehouse.gov/share/brain-initiative). Even when calculating the mass of an isolated ion, the missing or additional electrons can generally be ignored, since their contribution to the overall mass is negligible, reflected only in the nonsignificant digits that will be lost when the computed mass is properly rounded. Understanding the relationship between the masses of atoms and the chemical formulas of compounds allows us to quantitatively describe the composition of substances. It is also important to understand the mechanism responsible for any changes in the number of neurotransmitter molecules released—for example, some dysfunction in exocytosis, a change in the number of vesicles in the neuron, or a change in the number of neurotransmitter molecules in each vesicle. It provides a specific measure of the number of atoms or molecules in a bulk sample of matter. The formula mass for an ionic compound is calculated in the same way as the formula mass for covalent compounds: by summing the average atomic masses of all the atoms in the compound's formula. (credit: modification of work by Mark Ott) Because the definitions of both the mole and the atomic mass unit are based on the same reference substance, 12C, the molar mass of any substance is numerically equivalent to its atomic or formula weight in amu. 9.545 × 10²² molecules C₄H₁₀; 9.545 × 10²³ atoms H The brain is the control center of the central nervous system (Figure 8). The recommended daily dietary allowance of vitamin C for children aged 4–8 years is 1.42 × 10⁻⁴ mol. We can argue that modern chemical science began when scientists started exploring the quantitative as well as the qualitative aspects of chemistry. Figure 9. What is the formula mass (amu) of calcium phosphate? Clockwise from the upper left: 130.2 g of C₈H₁₇O₇ (1-octanol), formula mass 130.2 amu, 454.4 g of HgI₂ (mercury(II) iodide, formula mass 454.4 amu), 32.0 g of CH₃OH (methanol, formula mass 32.0 amu) and 256.5 g of S₈ (sulfur, formula mass 256.5 amu). Table salt, NaCl, contains an array of sodium and chloride ions combined in a 1:1 ratio. Carrying out the two-step computation yields: $(1.47 \text{ g}) / (0.5 \text{ g/mol}) = 2.94 \text{ mol}$. The factor-label method yields the desired cancellation of units, and the computed result is on the order of 10²² as expected. These studies also indicate that not all of the dopamine in a given vesicle is released during exocytosis, suggesting that it may be possible to regulate the fraction released using pharmaceutical therapies. (b) Information is transmitted in brain tissue and throughout the central nervous system by specialized cells called neurons (micrograph shows cells at 1600× magnification). These neurotransmitter molecules are stored in small intracellular structures called vesicles that fuse to the cell wall and then break open to release their contents when the neuron is appropriately stimulated. Likewise, if we know the number of moles of a substance, we can derive the number of atoms or molecules and calculate the substance's mass. Its formula mass is 58.44 amu. Figure 3. Watch this video and then complete the "Think" questions that follow. The molar mass of an element (or compound) is the mass in grams of 1 mole of that substance, a property expressed in units of grams per mole (g/mol) (see Figure 4). What is the estimated average requirement of potassium in moles? Specialized cells called neurons transmit information between different parts of the central nervous system by way of electrical and chemical signals. (a) Chemical signals are transmitted from neurons to other cells by the release of neurotransmitter molecules into the small gaps (synapses) between the cells. Deriving Moles from Grams for a Compound Our bodies synthesize protein from amino acids. Deriving the Number of Atoms and Molecules from the Mass of a Compound A packet of an artificial sweetener contains 40.0 mg of saccharin (C₇H₅NO₃S), which has the structural formula: Given that saccharin has a molar mass of 183.18 g/mol, how many saccharin molecules are in a 40.0-mg (0.0400-g) sample of saccharin? How many moles of glycine molecules are contained in 28.35 g of glycine? (b) Dopamine, C₈H₁₁NO₂, is a neurotransmitter involved in a number of neurological processes. This process is called exocytosis (see Figure 9). The model shows the molecular structure of aspirin, C₉H₈O₄. Each sample contains 6.02 × 10²³ molecules or formula units—1.00 mol of the compound or element. One neurotransmitter that has been very extensively studied is dopamine, C₈H₁₁NO₂. In an earlier chapter, we described the development of the atomic mass unit, the concept of average atomic masses, and the use of chemical formulas to represent the elemental makeup of substances. This experimental approach required the introduction of a new unit for amount of substances, the mole, which remains indispensable in modern chemical science. What is the formula mass (amu) of this compound? The given mass of K (4.7 g) is a bit more than one-tenth the molar mass (39.10 g), so a reasonable "ballpark" estimate of the number of moles would be slightly greater than 0.1

mol. Following the approach outlined above, the formula mass for this compound is calculated as follows: Check Your Learning Calcium phosphate, $\text{Ca}_3(\text{PO}_4)_2$, is an ionic compound and a common anti-caking agent added to food products. Check Your Learning What is the mass of 0.443 mol of hydrazine, N_2H_4 ? Figure 4. Computing Molecular Mass for a Covalent Compound Ibuprofen, $\text{C}_{13}\text{H}_{18}\text{O}_2$, is a covalent compound and the active ingredient in several popular nonprescription pain medications, such as Advil and Motrin.

6/5/2022 · Jun 18, 2020 · Amanda Nicole "Amethyst" Tarver Turner is an educator, Pagan minister, rootworker, spiritual adviser, life coach, end-of-life doula, and reiki/chakra healer. 2016. 2. Her practice is heavily influenced by hoodoo, Appalachian folk magic, and ... Password requirements: 6 to 30 characters long; ASCII characters only (characters found on a standard US keyboard); must contain at least 4 different symbols; BibMe Free Bibliography & Citation Maker - MLA, APA, Chicago, Harvard Fully Responsive Design. Your website will be displayed perfect on desktop, tablet and mobile devices. 6/5/2022 · Inverse Functions Practice Worksheet Algebra 2 May 6, 2022 May 6, 2022 1 3 2 3 5 1 6 4 2. 5 7 6 8 1 2 10 3 finding inverses find an equation for the inverse... We have writers who are well trained and experienced in different writing and referencing formats. Are you having problems with citing sources? Achiever Papers is here to help you with citations and referencing. Our professional writers are experienced in all formatting styles such as APA, MLA, Chicago, Turabian, and others. hace 2 dias · Check table J and see if the single element is above the element it will replace based on activity (metals replace metals and nonmetals replace nonmetals). Using thermometers, cotton balls, string and water, students make simple psychrometers—a tool that measures humidity. Download File PDF Mole Ratio Worksheet Answer Key.

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